

Highly Stable and Efficient Light-Emitting Electrochemical Cells Based on Cationic Iridium Complexes Bearing Arylazole Ancillary Ligands

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A series of bis-cyclometalated iridium(III) complexes of general formula $[\text{Ir}(\text{ppy})_2(\text{N}^{\wedge}\text{N})][\text{PF}_6]$ ($\text{ppy}^- = 2\text{-phenylpyridinate}$; $\text{N}^{\wedge}\text{N} = 2\text{-(1H-imidazol-2-yl)pyridine}$ (**1**), $2\text{-(2-pyridyl)benzimidazole}$ (**2**), $1\text{-methyl-2-pyridin-2-yl-1H-benzimidazole}$ (**3**), $2\text{-(4'-thiazolyl)benzimidazole}$ (**4**), and $1\text{-methyl-2-(4'-thiazolyl)benzimidazole}$ (**5**)) is reported, and their use as electroluminescent materials in light-emitting electrochemical cell (LEC) devices is investigated. **[2][PF₆]** and **[3][PF₆]** are orange emitters with intense unstructured emission around 590 nm in acetonitrile solution. **[1][PF₆]**, **[4][PF₆]**, and **[5][PF₆]** are green weak emitters with structured emission bands peaking around 500 nm. The different photophysical properties are due to the effect that the structure of the ancillary ligand has on the nature of the emitting triplet state. Whereas the benzimidazole unit stabilizes the LUMO and gives rise to a ³MLCT/³LLCT emitting triplet in **[2][PF₆]** and **[3][PF₆]**, the presence of the thiazolyl ring produces the opposite effect in **[4][PF₆]** and **[5][PF₆]** and the emitting state has a predominant ³LC character. Complexes with ³MLCT/³LLCT emitting triplets give rise to LEC devices with luminance values one order higher than complexes with ³LC emitting states. Protecting the imidazole N–H bond with a methyl group, as in complexes **[3][PF₆]** and **[5][PF₆]**, determine that the emissive properties become more stable. **[3][PF₆]** leads to outstanding LECs with simultaneously high luminance (904 cd m⁻²), efficiency (9.15 cd A⁻¹), and stability (lifetime over 2500 hours).

INTRODUCTION

Energy efficient devices for lighting and display applications are highly demanded. During the last years, organic light-emitting diodes (OLEDs)^{1,2,3} have been widely used in energy efficient displays, but their high manufacturing costs make them not suitable for a wide use in lighting. Light-emitting electrochemical cells (LECs)^{4,5} have emerged as potential candidates for light-emitting devices with promising energy savings and high-quality lighting. LECs are composed of a single active layer of an ionic organic or organometallic semi-

conductor placed between two electrodes. Due to their simple single-layer structure and their easy processing from solution, LECs are compatible with low-cost manufacturing⁶ and constitute a good alternative to OLEDs for lighting. A wide range of organic materials has been used in the manufacture of LECs.^{7,8,9,10,11} However, the best efficiencies have been achieved with ionic transition-metal complexes (iTMCs),^{12,13,14,15} and, in particular, with iridium(III) complexes due to their efficient spin-orbit coupling, which ideally allows to harvest 100% of the excitons generated in the device.¹⁶

The optical and emitting properties of Ir-iTMCs can be easily modulated by chemically modifying the structure of the ligands attached to the central Ir(III) core. Nowadays, a wide variety of complexes has been designed and used to manufacture LEC devices covering the whole visible spectrum,^{16,17,18,19} with high efficiencies,^{20,21,22} top lifetimes,^{23,24,25} and short turn-on times.^{26,27,28} However, it is still to be found a complex that allows the fabrication of a device that reunites all those characteristics and becomes suitable for real-life lighting applications.

One desirable key feature for LEC devices is a long lifetime, which should be in the range of a few thousands hours if LECs have to compete with commonly used lighting devices. A long device lifetime is related to a high complex stability. The most common strategy to achieve such high stability has been the inclusion of bulky hydrophobic ligands to protect the coordination center.^{23,29} However, bulky groups are known to slow down the migration of the ions in the device inducing long turn-on times.^{30,31} Hence, there is still room to explore new strategies and ligands in search for higher performances and stabilities.

In this work, we introduce a series of new complexes with arylazole-based ancillary ligands without bulky groups, and use them to prepare LEC devices trying to reach high stabilities without affecting the turn-on times. Arylazoles present some advantages as ancillary ligands to prepare iridium iTMCs. They are either commercially available or easily prepared and offer a large deal of structural and chemical diversity, since a variety of aryl and azole entities (imidazole, oxazole, thiazole, etc.) can be used. In addition, the enhanced electron-donating nature of the azole rings, due to the presence of one additional heteroatom with one or two lone electron pairs (N–H, O, or S), tends to destabilize the lowest-unoccupied molecular orbital (LUMO), enlarging the optical gap and blue-shifting the light emitted in comparison to complexes bearing classical 2,2'-bipyridine (bpy)-type ligands. Arylazole ligands can therefore be used to fine tuning the photophysical and chemical properties of Ir-iTMCs.³² In particular, the imidazole ring can furthermore be functionalized through the reactive N–H bond.

The electroluminescence properties of Ir-iTMCs bearing arylazole units as cyclometalating ligands have been widely investigated in LECs.^{33,34,35,36,37} In contrast, few examples of Ir-iTMCs with arylazole-based ancillary ligands have been tested in LECs. He et al.³⁸ described LEC devices using a family of Ir-iTMCs incorporating functionalized imidazole-type ancillary ligands with peak external quantum efficiencies up to 6.1%, a maximum luminance of 395 cd m⁻², and different emission colors depending on the complex. One year later, the same group reported a blue-green LEC with a pyrazole-type ancillary ligand and a peak external quantum efficiency of 7.6%, although at a much lower luminance, less than 20 cd m⁻².³⁹ In both cases, the devices exhibited long turn-on times from several minutes to hours and a poor device stability under constant voltage operation. More recently, Choe and col. have reported and characterized LECs based on aryl-imidazole ancillary ligands exhibiting high brightness and color tunability, although no study of the time-response characteristics is reported.^{40,41,42,43}

In this work, the complexes [1][PF₆]-[5][PF₆] shown in Figure 1 have been synthesized, characterized, and employed to develop new LEC devices. Their general formula is

[Ir(C[^]N)₂(N[^]N)][PF₆], where the C[^]N ligand always corresponds to 2-phenylpyridinate (ppy⁻) and the ancillary N[^]N ligand is 2-(1H-imidazol-2-yl)pyridine (pyim), 2-(2-pyridyl)benzimidazole (pybim), 1-methyl-2-pyridin-2-yl-1H-benzimidazole (pyMebim), 2-(4'-thiazolyl)benzimidazole (thiabenzazole, tbz), and 1-methyl-2-(4'-thiazolyl)benzimidazole (Metbz). The complexes were selected to study: i) the importance that the protection of the N–H bond of the imidazole has on the stability of the complexes, and ii) the influence that the substitution of the pyridyl group by the more electron-donating thiazolyl group has on the photophysical properties. The effect of the ancillary ligand on the electrochemical and photophysical properties was investigated and further analyzed with the help of density functional theory (DFT) calculations. The electroluminescence properties of the complexes were investigated in LEC devices operated under pulsed driving conditions. Complex [2][PF₆] was previously known,⁴⁴ but it has not been used in LEC devices yet. Complex [3][PF₆] and its use in LECs under voltage scan (current density and luminance) were reported but, to the best of our knowledge, no investigation of the device performance vs. time was performed.⁴⁰

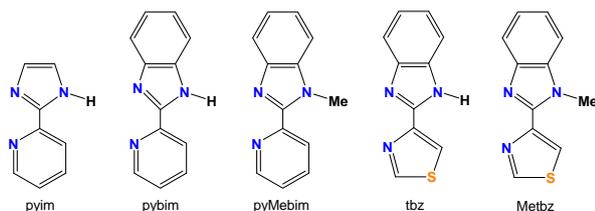
RESULTS AND DISCUSSION

Synthesis and Characterization

The complexes depicted in Figure 1 were prepared following the Nonoyama protocol⁴⁵ in two steps. In the first step, the chloro-bridged dimeric iridium complex [Ir(μ-Cl)(ppy)₂]₂ was isolated from the reaction between IrCl₃·xH₂O and 2-phenylpyridine (Hppy). In the second step, the corresponding ancillary N[^]N ligand was introduced in the coordination sphere of Ir(III) through a bridge splitting reaction, and then, an anion exchange protocol was applied to isolate the PF₆⁻ salt, as outlined in Scheme 1. The complexes were isolated in good yields as racemic mixtures (Δ, Λ) of yellow-to-orange solids that are soluble in common organic solvents.

All the complexes were characterized by multinuclear NMR, IR, and mass spectroscopy and by elemental analysis. 2D NMR experiments were conducted to assign all the signals in the ¹H (Figure S1 and S2 in the Supporting Information) and ¹³C{¹H} NMR spectra.

Ancillary ligands (N^N)



Complexes [1][PF₆]-[5][PF₆]

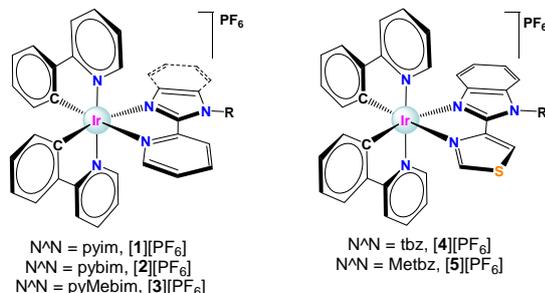
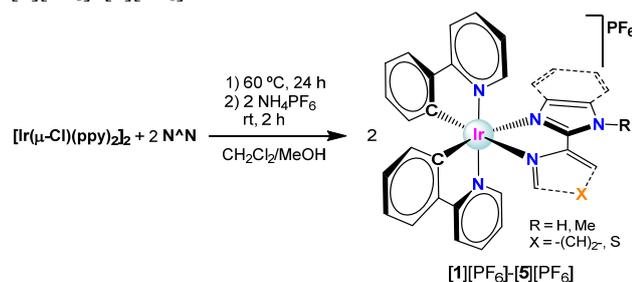


Figure 1. Arylazole ligands used in this work and chemical structures of the synthesized iridium(III) complexes [1][PF₆]-[5][PF₆].

Scheme 1. Second step in the synthetic route for complexes [1][PF₆]-[5][PF₆].



The following main features are inferred from the analysis of the NMR spectra recorded in DMSO-*d*₆ at 25 °C:

(1) The asymmetric nature of the synthesized compounds (C₁-symmetry) renders the two phenylpyridinate ligands in every complex inequivalent and thus the corresponding ¹H and ¹³C{¹H} NMR spectra show two different sets of signals for the C^N donors.

(2) The 2D noesy spectra of some complexes exhibit inter-ligand and inter-ring crosspeaks very useful in the assignment of all the signals (Figure S2).

(3) Complexes [1][PF₆], [2][PF₆], and [4][PF₆] exhibit a broad peak between 14 and 15 ppm for its N–H proton, whereas complexes with N–Me groups display a sharp singlet at 4.45 or 4.31 ppm for [3][PF₆] and [5][PF₆], respectively.

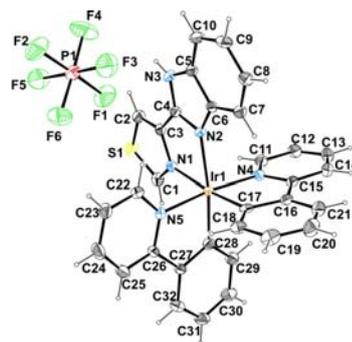
The ¹³C{¹H} NMR spectra are fully consistent with the structural features ascertained by ¹H NMR. In the FAB+ MS spectra, the characteristic sets of peaks for the cationic fragments [Ir(C^N)₂(N^N)]⁺ and [Ir(C^N)₂]⁺ are found, with isotopic patterns compatible with the presence of iridium, whereas in the FT-IR spectra typical diagnostic bands are observed for the PF₆⁻ counteranion (see the Supporting Information).

The thermal stability of complexes [1][PF₆]-[5][PF₆] was investigated by thermogravimetric analysis (TGA) under nitrogen flow. All the complexes are thermally stable with decomposition temperatures (*T*_d, corresponding to 5% weight

loss) all over 300 °C (Figure S3). Interestingly, complexes with N–Me groups are more stable than those with N–H groups (for instance, *T*_d = 306 and 386 °C for [2][PF₆] and [3][PF₆], respectively).

Crystal Structures by X-ray Diffraction

Single crystals suitable for X-ray diffraction analysis were obtained for [4][PF₆] and [5][PF₆] by slow diffusion of *n*-hexane into a solution of [4][PF₆] in CH₃CN/CH₂Cl₂ and by slow evaporation of a solution of [5][PF₆] in CH₂Cl₂. Both complexes crystallize in the orthorhombic *Pbca* space group. The ORTEP diagrams for the asymmetric units are shown in Figure 2. Selected bond lengths and angles with estimated standard deviations are gathered in Table S1 in the Supporting Information, and relevant crystallographic parameters are given in Table S2. For both complexes, the corresponding unit cell consists of four pairs of the two possible enantiomers (Λ and Δ), which result from the helical chirality inherent to tris-bidentate octahedral metal complexes. Indeed, the cationic complexes adopt a pseudo octahedral geometry around the iridium center with the expected *cis*-C,C and *trans*-N,N disposition for the ppy⁻ ligands. The ancillary N^N ligands show N–Ir–N bite angles of 76.1(1)° (for tbz in [4][PF₆]) and 75.9(1)° (for Metbz in [5][PF₆]), in agreement with those previously reported for five-membered iridacycles in similar complexes.^{46,47} The cyclometalated C^N ligands have also standard values (~80°) for the C–Ir–N bite angles. In both compounds, the 3D crystal architecture is built on the foundation of hydrogen bonding contacts and anion-π interactions in which the PF₆⁻ anions occupy a central position (Figure S4). Furthermore, both crystal structures show dimeric entities derived from π–π stacking interactions involving the benzimidazole units of two vicinal cation complexes. Conspicuously, these π–π interactions are hampered by the presence of the N–Me group in [5][PF₆], and hence they are very weak relative to the analogous interactions in [4][PF₆], where the centroid-centroid distance is shorter, 3.58 vs 3.87 Å, (see Table S3 and Figures S5 and S6).



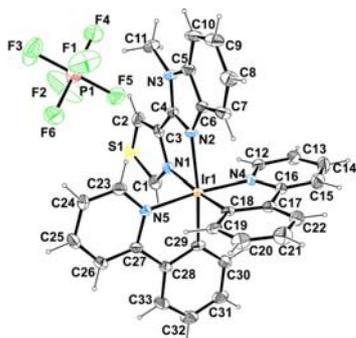


Figure 2. ORTEP diagrams for the asymmetric unit of compounds (Λ)-[4][PF₆] (top) and (Λ)-[5][PF₆] (bottom). Thermal ellipsoids are shown at the 30% probability.

Electrochemical behavior

The redox potentials of complexes [1][PF₆]–[5][PF₆] were experimentally determined in acetonitrile solution by cyclic voltammetry using [tⁿBu₄N][PF₆] as supporting electrolyte and glassy carbon as working electrode. Potentials were defined with respect to the ferrocene/ferrocenium (Fc/Fc⁺) couple. All the electrochemical experiments were performed under argon atmosphere. The cyclic voltammogram of every PF₆[−] salt shows a unique quasi-reversible oxidation peak in the anodic potential region between +0.74 and +0.79 V (see Table 1 and Figure 3), which is ascribed to the oxidation of the ppy-Ir(III) environment as discussed below.

The analysis of the cathodic region allowed to establish the following trends: (a) complexes with N–H groups in the ancillary ligand ([1][PF₆], [2][PF₆], and [4][PF₆]) show totally irreversible peaks at −2.20, −1.95, and −2.14 V, respectively, indicating the instability of the reduced species; (b) complex [3][PF₆], where the reactive N–H group of pybim has been replaced with a methyl group, displays a quasi-reversible reduction peak at −1.91 V, which supports the higher stability of [3][PF₆] versus [2][PF₆]; (c) complex [5][PF₆], also with a N–Me group, shows a reduction peak at −2.27 V giving rise to the largest electrochemical gap (3.04 V), and its irreversible nature suggests that the reduction mechanism is different to that of [3][PF₆]. These results anticipate a very good stability for complex [3][PF₆] under the operation conditions used in LEC devices owing to the reversible nature of the oxidation/reduction processes in which the complex is-involved in those devices.

Table 1. Cyclic Voltammetric Data Referred to Fc/Fc⁺ in Acetonitrile Solution (10^{−3} M).^a

| Compound | $E_{1/2}^{\text{ox}}$ (V) | $E_{1/2}^{\text{red}}$ (V) | $\Delta E_{1/2}$ (V) |
|-----------------------|---------------------------|----------------------------|----------------------|
| [1][PF ₆] | +0.74 (qr) | −2.20 (ir) | 2.94 |
| [2][PF ₆] | +0.77 (qr) | −1.95 (ir) | 2.72 |
| [3][PF ₆] | +0.79 (qr) | −1.91 (qr) | 2.70 |
| [4][PF ₆] | +0.76 (qr) | −2.14 (ir) | 2.90 |
| [5][PF ₆] | +0.77 (qr) | −2.27 (ir) | 3.04 |

^a Measured using 0.1 M [tⁿBu₄N][PF₆] as supporting electrolyte and a scan rate of 0.15 V s^{−1} (qr = quasi-reversible, ir = irreversible).

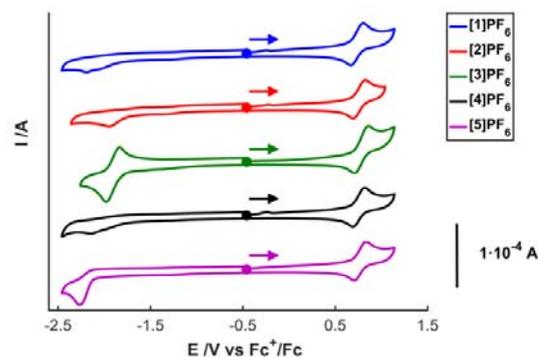


Figure 3. Cyclic voltammograms of complexes [1][PF₆]–[5][PF₆] in acetonitrile solution (10^{−3} M), using 0.1 M [tⁿBu₄N][PF₆] as supporting electrolyte. Scan rate of 0.15 V s^{−1}. (•) Indicates the starting (E_i) and final potential (E_f), E_i = E_f = −0.46 V (vs. Fc/Fc⁺), clockwise scan.

Photophysical properties

The UV-Vis absorption spectra recorded in acetonitrile solution are displayed in Figure 4. The complexes show intense absorption bands in the UV with maxima in the range 250–350 nm, which are mainly ascribed to spin-allowed $\pi \rightarrow \pi^*$ transitions of the ligands. Lower energy bands above 350 nm correspond to spin-allowed metal-to-ligand (¹MLCT) and ligand-to-ligand charge transfer (¹LLCT) excitations, whereas the low intensity tails above 450 nm arise from spin-forbidden ³MLCT, ³LLCT, and ligand-centered (³LC) transitions. Compared with complex [1][PF₆], taken as a reference, complexes [2][PF₆] and [3][PF₆] exhibit a more intense absorption centered around 325 nm due to the benzimidazole unit. Substitution of the pyridine by a thiazole ring in the C^N ligand of complexes [4][PF₆] and [5][PF₆] blue-shifts this absorption to 300 nm.

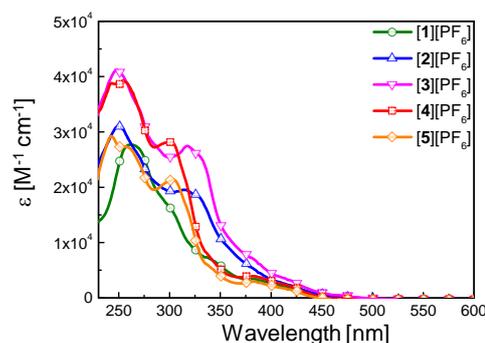


Figure 4. UV-vis absorption spectra of complexes [1][PF₆]–[5][PF₆] recorded in acetonitrile solution (5 × 10^{−5} M).

The photoluminescence properties of complexes [1][PF₆]–[5][PF₆] were investigated by excitation of both deoxygenated acetonitrile solutions (5 × 10^{−5} M, room temperature and 77 K) and thin films with the same composition used in the active layer of LEC devices (iTMC mixed with 1-butyl-3-methylimidazolium hexafluorophosphate in a 4:1 molar ratio).

Photoluminescence spectra are shown in Figure 5, and emission maxima wavelengths (λ_{em}), photoluminescence quantum

yields (Φ_{PL}), excited-state lifetimes (τ), and radiative (k_r) and non-radiative (k_{nr}) rate constants are summarized in Table 2. In solution, at room temperature, [2][PF₆] and [3][PF₆] show broad and unstructured emission bands, suggesting a large charge-transfer (CT) character for T₁ in these complexes. The CT character of the emitting state is furthermore supported by the hypsochromic shift observed in the emission bands of these complexes when the spectra are registered at 77 K (rigidochromic effect).^{48,49,50} In contrast, complexes [1][PF₆], [4][PF₆], and [5][PF₆] show emission profiles with some vibrational structure at room temperature that becomes more marked at 77 K and present no significant rigidochromic shift (Figure 5a and b). This behavior is typical of emissive states with a large ³LC character.^{51,52}

The emission color of the complexes in solution goes from yellow-orange ([2][PF₆] and [3][PF₆]) to blue-green ([1][PF₆], [4][PF₆], and [5][PF₆]), and the peak emission wavelengths follow the trend: [4][PF₆] (489, 508 nm) \approx [5][PF₆] (486, 508 nm) < [1][PF₆] (516 nm) < [2][PF₆] (583 nm) < [3][PF₆] (596 nm) (Table 2). Hence, on changing the ancillary ligand from pyim ([1][PF₆]) to pybim ([2][PF₆]), the emission experiences a red shift, whereas the replacement of pybim ([2][PF₆]) or pyMebim ([3][PF₆]) by tbz ([4][PF₆]) or Metbz ([5][PF₆]) leads to a blue shift comparable in magnitude. The substitution of the N-H group with a methyl in [3][PF₆] produces a small red shift (13 nm) on the emission peak respect to the non-substituted complex [2][PF₆] that is not observed when passing from [4][PF₆] to [5][PF₆].

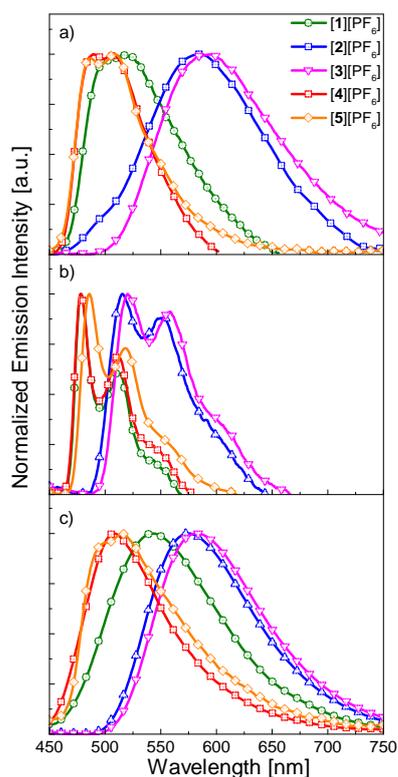


Figure 5. Photoluminescence spectra recorded at room temperature (a) and at 77 K (b) in deaerated acetonitrile solution (5×10^{-5} M) by excitation at 360 nm. (c) Photoluminescence spectra registered in thin film at room temperature by excitation at 320 nm.

Complexes [4][PF₆] and [5][PF₆] exhibit weak emission ($\Phi_{\text{PL}} = 9.0\%$ and 4.4% , respectively), in good agreement with the ³LC character predicted for the emissive state in these complexes. In comparison, complexes [2][PF₆] and [3][PF₆] feature significantly higher Φ_{PL} (38.5% and 39.4% , respectively) in accord with the expected ³MLCT character of the emissive state that enhances the spin-orbit coupling and therefore favors the intersystem crossing. The Φ_{PL} recorded for [3][PF₆] is similar to that previously reported by Sunesh et al.⁴⁰ The Φ_{PL} of complex [1][PF₆] (20.4%) is between both extremes, doubling the values registered for the LC-type complexes and being about one half of the values recorded for the MLCT-type complexes. This suggests that the electronic description of the emissive state for [1][PF₆] requires a deeper analysis.

Excited-state lifetimes measured in deoxygenated acetonitrile solution under an argon atmosphere are of the order of hundreds of nanoseconds (Table 2). The longest lifetime (485 ns) is recorded for complex [3][PF₆] and similar lifetimes around 200 ns are obtained for [2][PF₆], [4][PF₆] and [5][PF₆]. For [1][PF₆], a biexponential fit was used to determine an average lifetime value of 316 ns (see Table S4 in the Supporting Information for details). The lower Φ_{PL} and shorter τ values recorded for [4][PF₆] and [5][PF₆] determine the higher non-radiative decay rate constants obtained for these two complexes.

The emission spectra were also recorded for amorphous thin films of the complexes with the composition used in LEC devices (Figure 5c). The spectra show a similar shape for all the complexes, and the emission maxima follow the same trend observed in solution: [4][PF₆] (507 nm) < [5][PF₆] (514 nm) < [1][PF₆] (542 nm) < [2][PF₆] (574 nm) < [3][PF₆] (581 nm). Interestingly, complexes with CT emitting states ([2][PF₆] and [3][PF₆]) exhibit their emission maxima slightly shifted to shorter wavelengths compared to solution (Table 2), whereas those with LC emitting states ([4][PF₆] and [5][PF₆]) are red-shifted. [1][PF₆] undergoes the most drastic change in the emission maximum in passing from solution to thin film, pointing again to the existence of differences between the electronic nature of the emitting state in this complex compared with complexes [4][PF₆] and [5][PF₆]. The thin-film emission spectra present almost no vibrational structure, probably caused by the rigidity of the solid state, and only [5][PF₆] keeps observable vibrational bands.

Table 2. Photophysical properties of complexes [1][PF₆]–[5][PF₆] in Solution (5×10^{-5} M) and Thin Film.

| Complex | Solution (room temperature) | | | | | Solution (77 K) | Thin Film | |
|-----------------------|---|---------------------------------------|----------------------------|--|---|---|---|-------------------------------------|
| | λ_{em} (nm) ^{a,b} | Φ_{PL} (%) ^{a,b} | τ (ns) ^{a,c} | $k_r/10^5$ (s ⁻¹) ^d | $k_{\text{nr}}/10^5$ (s ⁻¹) | λ_{em} (nm) ^{a,b} | λ_{em} (nm) ^e | Φ_{PL} (%) ^e |
| [1][PF ₆] | 516 | 20.4 | 316 | 0.1 | 0.1 | 542 | 542 | 20.4 |
| [2][PF ₆] | 583 | 38.5 | 200 | 0.1 | 0.1 | 574 | 574 | 38.5 |
| [3][PF ₆] | 596 | 39.4 | 485 | 0.1 | 0.1 | 581 | 581 | 39.4 |
| [4][PF ₆] | 489 | 9.0 | 200 | 0.1 | 0.1 | 507 | 507 | 9.0 |
| [5][PF ₆] | 508 | 4.4 | 200 | 0.1 | 0.1 | 514 | 514 | 4.4 |

| | | | | | | | | |
|-----------------------|----------|------|------------------|------|------|----------|-------------|------|
| [1][PF ₆] | 516 | 20.4 | 316 ^f | 6.5 | 25.2 | 478, 511 | 542 | 5.5 |
| [2][PF ₆] | 583 | 38.5 | 222 | 17.3 | 27.7 | 515, 551 | 574 | 32.1 |
| [3][PF ₆] | 596 | 39.4 | 485 | 8.1 | 12.5 | 519, 557 | 581 | 44.6 |
| [4][PF ₆] | 489, 508 | 9.0 | 186 | 4.8 | 48.9 | 477, 512 | 507 | 3.3 |
| [5][PF ₆] | 486, 508 | 4.4 | 200 | 2.2 | 47.8 | 486, 518 | 495(sh) 514 | 6.0 |

^a Deoxygenated acetonitrile solution. ^b $\lambda_{\text{exc}} = 360$ nm. ^c $\lambda_{\text{exc}} = 365$ nm. $k_{\text{r}} = \Phi_{\text{PL}}/\tau$ and $k_{\text{t}} = (1 - \Phi_{\text{PL}})/\tau$. ^e $\lambda_{\text{exc}} = 320$ nm. ^f Average lifetime estimated using a biexponential fit (Table S4).

It is worth noting that, in contrast to the results observed in solution, where substitution of the N–H group with a methyl has no systematic effect, the presence of the methyl group produces an increase of the quantum yield in thin film. The Φ_{PL} augments from 32.1% in [2][PF₆] to 44.6% in [3][PF₆], and from 3.3% in [4][PF₆] to 6.0% in [5][PF₆]. The gain in Φ_{PL} is attributed to the fact that, as discussed above, the methyl group prevents the aggregation of the complex in the solid state and reduces self-quenching processes.

Theoretical calculations

Density functional theory (DFT) and time-dependent DFT (TD-DFT) calculations were performed on the cations of complexes [1][PF₆]⁺–[5][PF₆]⁺ to get a better understanding of the electrochemical and photophysical properties discussed above. Calculations were carried out at the B3LYP/(6-31G**+LANL2DZ) level including solvent effects (see the Experimental Section for details).

Calculations predict a near-octahedral structure for cations [1]⁺–[5]⁺ in their ground electronic state (*S*₀), in good agreement with the results observed for similar complexes in previous studies^{36,53,54,55,56} and the X-Ray structures reported above for complexes [4][PF₆]⁺ and [5][PF₆]⁺ (Figure 2). Both the cyclometalating and the ancillary ligands display coplanarity for their constituting rings, as it is usually found for iTMC complexes possessing five-atoms chelate rings and without bulky substituents in their molecular structure.³² The largest deviation from coplanarity is indeed found for the pyMebim N^N ligand (7.3°) in [3][PF₆]⁺ due to the methyl group introduced in the interannular region. The values predicted for the bite angles C–Ir–N (~80.1°) and N–Ir–N (~74.5° for [1]⁺, [2]⁺, and [4]⁺, and ~74.0° for [3]⁺ and [5]⁺) are very close to those obtained from the X-Ray structures of [4][PF₆]⁺ and [5][PF₆]⁺.

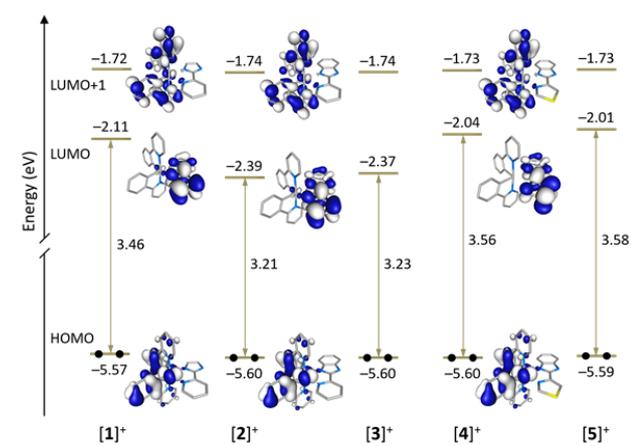


Figure 6. Energy diagram showing the isovalue contours (± 0.03 a.u.) and the energy values calculated for the HOMO, LUMO, and

LUMO+1 of complexes [1]⁺ to [5]⁺. The MOs of [3]⁺ and [5]⁺ are identical to those of [2]⁺ and [4]⁺ and are not displayed.

Figure 6 displays the isovalue contours calculated for the highest-occupied (HOMO) and lowest-unoccupied molecular orbital (LUMO) of cations [1]⁺–[5]⁺. The LUMO+1 is also displayed because it plays a relevant role in the photophysical properties (see below). As it is usually found for ppy-based cyclometalated Ir-iTMCs,^{5,54,57} the HOMO of cations [1]⁺–[5]⁺ results from a mixture of *d* _{π} orbitals of Ir(III) and phenyl π orbitals, with some contribution from the pyridine rings of the cyclometalating ligands. As the five complexes share the same C^N ligand, the HOMO has almost identical energies for all of them. This is in good agreement with the experimental $E_{1/2}^{\text{ox}}$ values (Table 1), which remain almost constant and imply the oxidation of the ppy-Ir environment. The same trend is found for the LUMO+1, which is centered on the cyclometalating ligands and presents almost the same energy for all the complexes (Figure 6).

In contrast to the HOMO, the LUMO is fully located over the ancillary N^N ligand and its energy varies upon modifying the identity of the ligand. The condensation of a benzene ring to the imidazole ring to produce a benzimidazole unit leads to the stabilization of the LUMO by ~0.3 eV due to the extended conjugation. On the contrary, the exchange of the pyridine ring by a thiazole ring in [4]⁺ and [5]⁺ compensates the effect of benzene condensation and leads to a strong destabilization of the LUMO (~0.35 eV) compared to [2]⁺ and [3]⁺. This destabilization is due to the strong electron-donating character of the thiazole ring resulting from the presence of the sulfur atom, which contributes with a lone electron pair to the aromatic π -system of the five-membered ring. Methylation of the imidazole ring has a very small effect, destabilizing the LUMO by around 0.02 eV. The changes predicted for the energy of the LUMO justify the less negative reduction potentials measured for [2][PF₆]⁺ (–1.95 V) and [3][PF₆]⁺ (–1.91 V) compared with [1][PF₆]⁺ (–2.20 V), and the increase of the potential in passing to [4][PF₆]⁺ (–2.14 V) and [5][PF₆]⁺ (–2.27 V). Upon reduction, the extra electron fully enters in the N^N ligand as illustrated by the unpaired-electron spin densities shown in Figure S7 for the radical species [3][•] and [5][•]. It is to be noted that for [5][•] a high spin density of 0.47e is located on the C2 atom of the thiazole ring (see Figure 2 for atom numbering). The highly localized nature of the unpaired electron could explain the instability of the reduced species of [5][PF₆]⁺ despite the N–H group of the pybim unit is protected with a methyl group as in [3][PF₆]⁺.

The size of the HOMO–LUMO energy gap is therefore mainly determined by the position of the LUMO and increases along the series [2]⁺ \approx [3]⁺ < [1]⁺ < [4]⁺ \approx [5]⁺. Obviously, if a direct correlation is established between the HOMO–LUMO gap and the emission energy, bluer emissions have to be expected along this series, which is indeed in agreement with the experimental trend observed for the emission maxima (Table 2).

However, the picture is not as simple since, as discussed in the previous section, experimental evidences point to a different electronic nature of the emitting triplet state. Therefore, we cannot directly presume that emission originates from the HOMO \rightarrow LUMO transition for all the complexes. TD-DFT calculations will help us to elucidate the nature of the emitting state and to rationalize the experimental observations.

TD-DFT calculations were performed for the lowest-energy triplet (T_n) excited states of cations $[1]^+$ to $[5]^+$. The results obtained for the four lowest triplets (vertical excitation energies, description in terms of orbital monoexcitations, and electronic nature) are detailed in Table S5 in the Supporting Information and are summarized in a graphical way in Figure 7. For each cation, two triplet states of LC nature with some MLCT character are found, which are mainly described by the HOMO \rightarrow LUMO+1 and HOMO \rightarrow LUMO+2 excitations. As all these orbitals are located on the C \wedge N ligands with some participation of the metal (Figure 6), they remain unchanged along the series and the resulting ${}^3\text{LC}(\text{C}\wedge\text{N})$ states present almost the same energy (2.75 and 2.79 eV) for all the complexes (Figure 7). The HOMO \rightarrow LUMO excitation implies an electron transfer from the ppy-Ir environment, where the HOMO is located, to the N \wedge N ligand, where the LUMO resides (Figure 6), and therefore gives rise to a triplet state of mixed ${}^3\text{MLCT}/{}^3\text{LLCT}$ character. The state is calculated as the lowest triplet at 2.73 eV for $[1]^+$, very close in energy to the ${}^3\text{LC}(\text{C}\wedge\text{N})$ states, and stabilizes to 2.51 eV for $[2]^+$ and $[3]^+$ (Figure 7), accordingly with the stabilization of the LUMO when pybim is used as ancillary ligand. For $[4]^+$ and $[5]^+$, the ${}^3\text{MLCT}/{}^3\text{LLCT}$ state increases in energy to 2.95 and 2.97 eV, respectively, because of the LUMO destabilization induced by the replacement of the pyridyl ring with the thiazole ring, and it is by no longer the lowest triplet. It appears as T_4 (Figure 7). A fourth state of mixed ${}^3\text{LC}(\text{N}\wedge\text{N})/{}^3\text{MLCT}/{}^3\text{LLCT}$ character mainly resulting from the HOMO-1 \rightarrow LUMO or HOMO-2 \rightarrow LUMO excitations is also computed within the low energy triplets.

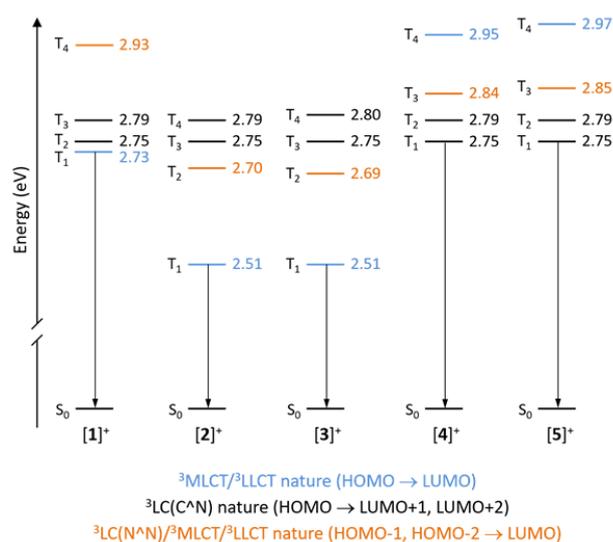


Figure 7. Energy diagram showing the energy values calculated for the lowest-energy triplet excited states (T_n) of complexes $[1]^+$ to $[5]^+$. The different electronic nature of the T_n states is denoted by using different colors.

The electronic nature of the lowest-lying T_1 triplet therefore changes along the series $[1]^+$ to $[5]^+$ (Figure 7). For $[2]^+$ and $[3]^+$, the HOMO \rightarrow LUMO ${}^3\text{MLCT}/{}^3\text{LLCT}$ state is lower in energy than the ${}^3\text{LC}(\text{C}\wedge\text{N})$ states and it clearly corresponds to the emitting state. This assignment supports the CT character proposed for this state in $[2][\text{PF}_6]$ and $[3][\text{PF}_6]$ on the basis of the broad and unstructured emission bands observed experimentally (Figure 5a). For $[4]^+$ and $[5]^+$, the destabilization of the ${}^3\text{MLCT}/{}^3\text{LLCT}$ state makes the ${}^3\text{LC}(\text{C}\wedge\text{N})$ triplets the lowest-lying states. The LC character of the emitting state explains the different photophysical behavior observed for complexes $[4][\text{PF}_6]$ and $[5][\text{PF}_6]$, which present structured emission bands with no significant rigidochromic shift at low temperatures (Figure 5a and b). For $[1]^+$, TD-DFT calculations predict that the ${}^3\text{MLCT}/{}^3\text{LLCT}$ state and the ${}^3\text{LC}(\text{C}\wedge\text{N})$ triplets are very close in energy and they can compete as emitting states. Experimental evidences actually point in this direction since the emission spectrum of $[1][\text{PF}_6]$ shows a nearly unnoticeable vibrational structure, and looks intermediate between those of $[2][\text{PF}_6]$ and $[3][\text{PF}_6]$ and those of $[4][\text{PF}_6]$ and $[5][\text{PF}_6]$ (Figure 5).

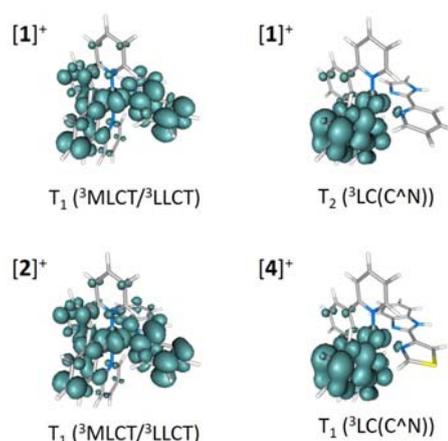


Figure 8. Unpaired-electron spin-density contours (0.002 a.u.) calculated for fully relaxed T_1 and T_2 states of $[1]^+$ and T_1 states of $[2]^+$ and $[4]^+$. The nature of the states is indicated within parentheses.

To get additional information about the emitting state, the geometries of the lowest triplet state of $[1]^+$ – $[5]^+$ were optimized using the spin-unrestricted UB3LYP approach. For $[1]^+$, both the $T_1({}^3\text{MLCT}/{}^3\text{LLCT})$ and $T_2({}^3\text{LC}(\text{C}\wedge\text{N}))$ triplets were calculated at the UB3LYP level. Figure 8 displays the unpaired-electron spin density distributions calculated for $[1]^+$ (T_1 and T_2), $[2]^+$, and $[4]^+$ at their optimized geometries. Those computed for $[3]^+$ and $[5]^+$ are identical to those shown for $[2]^+$ and $[4]^+$, respectively (Figure S8). For $[2]^+$ and $[3]^+$, T_1 features a spin density distribution spreading the ppy-Ir environment and the ligand (Ir 0.55e, C \wedge N ligands 0.43e, N \wedge N ligand 1.02e) that perfectly matches the topology of the HOMO \rightarrow LUMO MLCT/LLCT excitation. In contrast, for $[4]^+$ and $[5]^+$ the spin-density distribution is mainly centered on one of the cyclometalating ligands (~ 1.7 unpaired electrons) with a small contribution from the metal ($\sim 0.3e$). UB3LYP calculations therefore confirm the ${}^3\text{MLCT}/{}^3\text{LLCT}$ character of the emitting triplet for $[2]^+$ and $[3]^+$ and the predominant ${}^3\text{LC}$ nature of that triplet for $[4]^+$ and $[5]^+$. For $[1]^+$, the T_1 and T_2 states remain mostly isoenergetic after relaxation ($\Delta E = 2.59$ and 2.61 eV,

respectively, calculated as the difference between the total energies of S_0 and T_1/T_2 at their respective minimum energy structures), the ${}^3\text{MLCT}/{}^3\text{LLCT}$ state being slightly more stable. Thus, emission could take place both from T_1 or T_2 .

The calculated emission energies, estimated as the vertical energy difference between T_1 and S_0 at the optimized minimum-energy geometry of T_1 , follow closely the experimental values of the emission maxima. The values predicted for $[2]^+$ (590 nm) and $[3]^+$ (593 nm) are in good accordance with the experimental values (583 and 596 nm, respectively), whereas those obtained for $[1]^+$ (534 nm), $[4]^+$ (527 nm), and $[5]^+$ (528 nm) slightly overestimate the experimental values (516, 489–508, 486–508 nm, respectively).

Light-emitting electrochemical devices

Light-emitting electrochemical cells (LECs) were fabricated with complexes $[1][\text{PF}_6]$ – $[5][\text{PF}_6]$ in a double layer architecture consisting of a poly(3,4-ethylenedioxythiophene):poly(styrenesulfonate) (PEDOT:PSS) layer and the emissive layer sandwiched between indium tin oxide (ITO) and aluminum electrodes. The active layer contained the Ir-iTMC mixed with the ionic liquid (IL) 1-butyl-3-methylimidazolium hexafluorophosphate $[\text{Bmim}][\text{PF}_6]$ at a 4:1 (iTMC:IL) molar ratio. The IL was added to shorten the turn-on time of the LEC by increasing the concentration of ionic species and the ionic mobility in the light-emitting layer.^{26,58} After deposition of the emissive layer, the devices were transferred to an inert atmosphere glove box (<0.1 ppm O_2 and H_2O). The time-dependence of the luminance and average voltage of the LECs prepared was evaluated using a pulsed-current driving mode. LECs were operated by applying a block-wave pulsed current (frequency: 1 kHz, duty cycle: 50%), where the peak current density of the pulse was $200 \text{ A}\cdot\text{m}^{-2}$ and the average current density was $100 \text{ A}\cdot\text{m}^{-2}$. Pulsed-current driving has been demonstrated to lead to devices with longer lifetimes and faster response compared to fixed-current/voltage driving.²³

The dependence of the luminance and average voltage with time is depicted in Figure 9 and key parameters derived from the graphs are collected in Table 3. All the devices present a fast initial decrease of the voltage until a minimum stable value (2.3–2.7 V) is reached. This behavior is typical under the pulsed-current driving applied due to the operation mechanism of LECs. The voltage required to maintain the current density decreases versus time due to the formation of p- and n-doped regions, which reduces the resistance of the active layer. The fact that the voltage remains in all cases close to the minimum along the device operation time reveals that there are no signs of charge transport issues or chemical degradation. Except for device with $[4][\text{PF}_6]$, that does not emit any light, the luminance increases until a maximum is reached and then starts to decrease. The parameters used to quantify the response of the device are the time to reach this maximum luminance (t_{max}) and the time to reach a luminance value of $100 \text{ cd}\cdot\text{m}^{-2}$ (t_{on}). The t_{on} for devices with $[1][\text{PF}_6]$ and $[5][\text{PF}_6]$

was not established because these LECs did not achieve the threshold of $100 \text{ cd}\cdot\text{m}^{-2}$, yet they started up very rapidly and reached luminance levels of 18 and $51 \text{ cd}\cdot\text{m}^{-2}$ within 4 s, respectively.

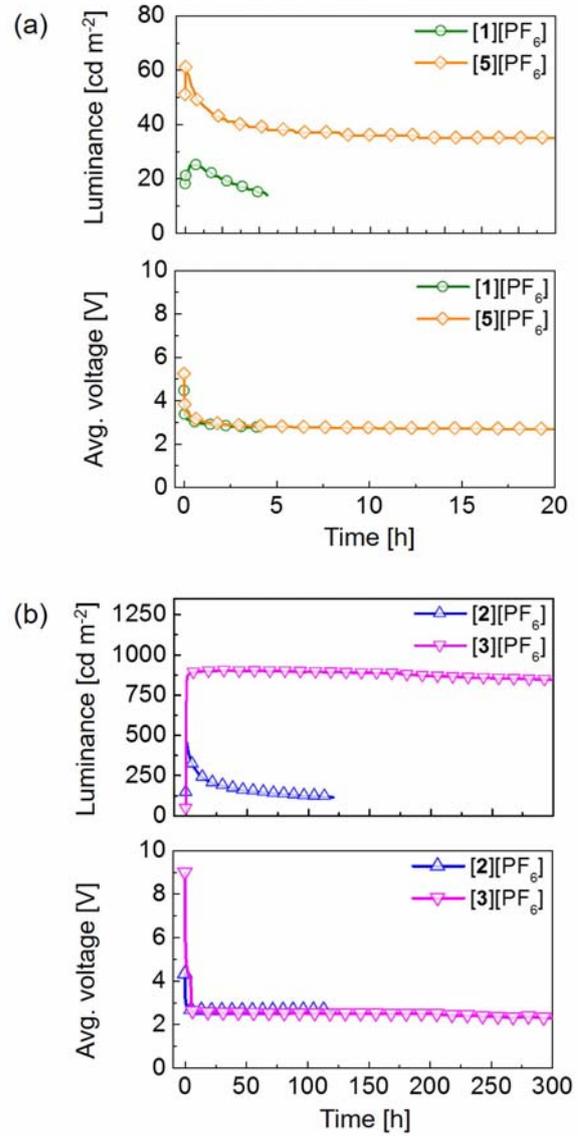


Figure 9. Luminance and average voltage for ITO/PEDOT:PSS/iTMC:[Bmim][PF₆] (4:1 molar ratio)/Al LEC devices measured by applying a block-wave pulsed current of $100 \text{ A}\cdot\text{m}^{-2}$ at a frequency of 1 kHz and duty cycles of 50%. a) iTMC = $[1][\text{PF}_6]$ and $[5][\text{PF}_6]$; b) iTMC = $[2][\text{PF}_6]$ and $[3][\text{PF}_6]$.

Table 3. Performance Parameters Obtained for ITO/PEDOT:PSS/Active Layer/Al LEC Devices Measured by Applying a Block-Wave Pulsed Current of $100 \text{ A}\cdot\text{m}^{-2}$ at a Frequency of 1 kHz and Duty Cycles of 50%. Active Layer: $[1][\text{PF}_6]$, $[2][\text{PF}_6]$, $[3][\text{PF}_6]$, and $[5][\text{PF}_6]$: $[\text{Bmim}][\text{PF}_6]$ at a 4 : 1 Molar Ratio.

| Complex | Lum ₀ ^a (cd·m ⁻²) | Lum _{max} ^b (cd·m ⁻²) | t _{on} ^c (s) | t _{max} ^d (min) | t _{1/2} ^e (h) | Efficacy ^f (cd·A ⁻¹) | EQE ^g (%) | PCE ^h (lm·W ⁻¹) |
|---------|--|--|-------------------------------------|--|--------------------------------------|--|-------------------------|---|
|---------|--|--|-------------------------------------|--|--------------------------------------|--|-------------------------|---|

| | | | | | | | | |
|-----------------------|-----|-----|-----|------|-------|-----|-----|-----|
| [1][PF ₆] | 18 | 25 | - | 15 | 2.25 | 0.3 | 0.1 | 1.3 |
| [2][PF ₆] | 143 | 460 | < 4 | 18 | 0.30 | 4.6 | 1.4 | 2.5 |
| [3][PF ₆] | 44 | 904 | 6.9 | 1680 | 2700* | 9.2 | 3.0 | 5.7 |
| [5][PF ₆] | 51 | 61 | - | 1 | 200 | 0.6 | 0.2 | 0.3 |

^aInitial luminance. ^bMaximum luminance reached. ^cTime to reach 100 cd·m⁻² luminance. ^dTime to reach the maximum luminance. ^eTime to reach one-half of the maximum luminance after t_{\max} is attained. ^fMaximum efficacy (ratio luminance/average current). ^gMaximum external quantum efficiency reached. ^hMaximum power conversion efficiency reached. *Extrapolated value.

The absence of light emission for [4][PF₆] and the low luminance observed for [1][PF₆] and [5][PF₆] can be related with the nature of the emitting state and, in the case of [4][PF₆], and probably [1][PF₆] as well, with the observed formation of dimeric entities in the crystal that favor self-quenching processes. As discussed above, the emitting state of these complexes has a predominant ³LC character with small participation of the iridium center. This limits the intersystem crossing and, as a consequence, low Φ_{PL} values are obtained for [1][PF₆] (5.5%), [4][PF₆] (3.3%), and [5][PF₆] (6.0%) in thin film (Table 2). In contrast, [2][PF₆] and [3][PF₆] have emitting states of ³MLCT/³LLCT nature with higher Φ_{PL} values (32.1 and 44.6%, respectively), and their LECs feature much higher luminances (Lum_{max} = 460 and 906 cd m⁻², respectively), comparable to those of the devices with the highest luminance maxima available in the literature for this wavelength range.^{24,11} The higher luminance registered for the [3][PF₆] device combines with a top efficiency (efficacy of 9.2 cd A⁻¹ and external quantum efficiency of 3.0%), and is in good agreement with the higher Φ_{PL} recorded for this complex in thin film. In addition, the LECs with [2][PF₆] and [3][PF₆] show a fast response, the longest t_{on} (7 s) being found for the LEC with [3][PF₆].

The stability of the devices was estimated as the decay time required for reaching one-half of the maximum luminance ($t_{1/2}$) after the t_{\max} is attained. The values obtained for $t_{1/2}$ (Table 3) indicate that the condensation of a benzene to the imidazole ring, which leads to increased luminescence, does not improve the device lifetime. In fact, the device made with the pybim-containing complex [2][PF₆] features a short $t_{1/2}$ of only 0.30 h, shorter than that of the pyim-containing [1][PF₆] device (2.25 h). These low device stabilities are in line with the instability of the reduced species observed electrochemically, which is attributed to the high reactivity of the N–H bond. The presence of this bond therefore is likely to be the origin of the limited device stability of [1][PF₆], [2][PF₆], and also [4][PF₆] which emits no light. However, when the imidazole N–H bond is substituted with a methyl group, the LEC stability improves dramatically, and $t_{1/2}$ increases up to 200 h in [5][PF₆] and above 2500 h in [3][PF₆]. The luminance of the [3][PF₆] device decreases only by 12% after 650 h (27 days) of continuous operation, and the linear extrapolation of the time dependence of luminance provides the value of $t_{1/2} = 2700$ h, which is a remarkable result for a LEC (Figure 9). The value of $t_{1/2}$ extrapolated for [3][PF₆] is among the best values available in the bibliography for LECs. In a recent review,¹¹ Henwood and Zysman-Colman concluded that the most stable emitter reported to date under the same pulsed current operation of 100 A m⁻² was presented by Tordera et al.,⁵⁹ who described a LEC device based on an imidazole-including complex ([Ir(ppy)₂(imp)] (imp = 1H-imidazo[4,5-f][1,10]phenantroline)) that featured high lifetime and luminescence and low turn-on times. The results obtained here for

the device manufactured with complex [3][PF₆] are even better than those of the device with [Ir(ppy)₂(imp)][PF₆]. The device stability obtained when using [3][PF₆] is longer ($t_{1/2} = 2700$ vs. 2000 h), the turn-on time is shorter (6.9 vs. 45 s), the maximum luminance is higher (904 vs. 684 cd m⁻²), and the maximum current efficiency is also higher (9.2 vs 6.5 cd A⁻¹). The values obtained here are also in excess of the most stable yellow-orange emitting LEC ($t_{1/2} \approx 4000$ h, 3.6 cd A⁻¹), that was operated under different pulsed current conditions (30% of duty cycle).²³ They are also competitive with the top peak efficiencies (18.6 cd A⁻¹) obtained by He et al. under constant voltage operation that leads to higher peak luminances and higher currents accompanied by a strong reduction in lifetime.³⁸ Interestingly, the exceptional efficiency found for the [3][PF₆] LEC (9.2 cd A⁻¹) is much higher than the efficiency previously obtained for this complex under voltage scan (1.5 cd A⁻¹).⁴⁰

Figure 10 shows the electroluminescence (EL) spectra of the four LECs and Table 4 summarizes the key parameters obtained from the EL spectra. The position of the emission maxima is similar to that recorded in the thin-film photoluminescence spectra for [1][PF₆], [2][PF₆], and [3][PF₆] (Table 2 and 4). However, the emission is red-shifted by 47 nm in the device of complex [5][PF₆], which presents a shoulder at 493 nm that was also observed in the PL spectrum in thin film (Table 2). The shift of the emission could be due to the electric field present in the device that stabilizes the ³LC emitting state of the complex, whereas the shoulder would correspond to the emission from a close triplet state (Table S5) not affected by the electric field. LEC devices emit green light for [1][PF₆] and [5][PF₆] and yellow light for [2][PF₆] and [3][PF₆].

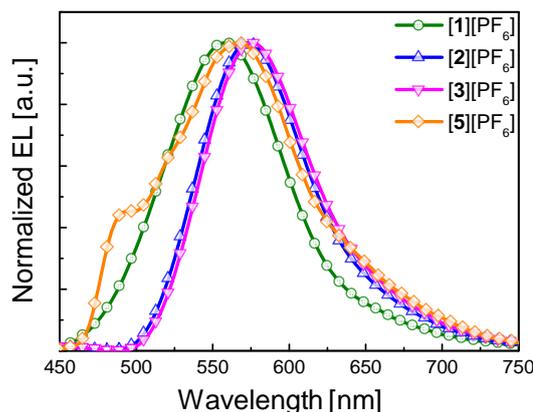


Figure 10. Normalized electroluminescence (EL) spectra of LECs based on complexes [1][PF₆], [2][PF₆], [3][PF₆], and [5][PF₆].

Table 4. Electroluminescence Parameters of LECs under Block-Wave Pulsed Current of 100 A m⁻² at a Frequency of 1 kHz and Duty Cycles of 50%.

| Complex | $\lambda_{\text{max,EL}}$ (nm) | CIE coordinates |
|-----------------------|--------------------------------|------------------|
| [1][PF ₆] | 558 | (0.4008, 0.5570) |
| [2][PF ₆] | 573 | (0.4811, 0.5109) |
| [3][PF ₆] | 576 | (0.4933, 0.5007) |
| [5][PF ₆] | 567 (493 sh) | (0.4138, 0.5219) |

CONCLUSIONS

A series of [Ir(ppy)₂(N[^]N)][PF₆] complexes with 2-phenylpyridinate cyclometalating ligands and different N[^]N arylazole ancillary ligands has been prepared and fully characterized. X-ray structural data are reported for complexes [4][PF₆] and [5][PF₆] incorporating thiazolyl rings.

The complexes show significantly different photoluminescence properties in solution depending on the chemical structure of the ancillary ligand. Complexes [2][PF₆] and [3][PF₆], bearing 2-(2-pyridyl)benzimidazole (pybim) and 1-methyl-2-pyridin-2-yl-1H-benzimidazole (pyMebim) N[^]N ligands, respectively, exhibit broad unstructured emission bands with maxima around 590 nm and experience a notable rigidochromic effect when recorded at 77 K. In contrast, complexes [4][PF₆] and [5][PF₆], in which the pyridyl ring of the N[^]N ligand is substituted by a thiazolyl ring, present vibrationally-structured emission bands peaking around 500 nm with no significant rigidochromic shift. DFT calculations fully justify the different photophysical behavior. The presence of the benzimidazole unit in [2][PF₆] and [3][PF₆] strongly stabilizes the LUMO (located on the ancillary ligand) and determines a ³MLCT/³LLCT character for the lowest-energy emitting triplet state. In contrast, the incorporation of the thiazolyl ring destabilizes the LUMO and thereby the HOMO → LUMO ³MLCT/³LLCT state in [4][PF₆] and [5][PF₆], and the emitting triplet has a predominant ³LC nature. The different nature of T₁ explains the low PLQYs (<10%) measured for [4][PF₆] and [5][PF₆]. Complex [1][PF₆], bearing a 2-(1H-imidazol-2-yl)pyridine N[^]N ligand, displays an intermediate behavior due to close energy of the ³MLCT/³LLCT and ³LC states.

Light-emitting electrochemical cells prepared with complexes [1][PF₆]–[5][PF₆] also present very different performances, in line with the photophysical and electronic structure differences mentioned above. LECs prepared with [2][PF₆] and [3][PF₆] (³MLCT/³LLCT emissive state) feature yellow color, high luminances (460 and 904 cd m⁻²) and efficiencies (4.61 and 9.15 cd A⁻¹), and low turn-on times (a few seconds). In contrast, LECs prepared with complexes [1][PF₆] and [5][PF₆], for which the emissive state has an important ³LC character, exhibit green color and low luminances (25 and 61 cd m⁻²). LECs with complexes bearing unprotected imidazole N–H bonds show short lifetimes ([1][PF₆] and [2][PF₆]) or emit no light ([4][PF₆]). The LEC based on [3][PF₆] exhibits record performance, simultaneously presenting high luminance and efficiency, short turn-on time, and high stability reaching lifetimes over 2500 hours.

The use of the benzimidazole unit in the ancillary ligand without incorporating bulky groups is therefore confirmed as an efficient strategy to obtain LECs with long lifetimes, short turn

on times, and high luminances, provided that the N–H group of the imidazole ring is conveniently protected with a methyl substitution.

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EXPERIMENTAL SECTION

General Methods and starting materials.

Starting materials. IrCl₃·xH₂O was purchased from Johnson Matthey and used as received. [Ir(μ-Cl)(ppy)₂]₂ (Hppy = phenylpyridine) was prepared according to literature procedures.^{45,60,61} The salts AgOTf, [NH₄][PF₆], and NaBPh₄ and the ligands 2-(2-pyridyl)benzimidazole (pybim) and 2-(4'-thiazolyl)benzimidazole (thiabenzazole, tbz) were purchased from Sigma-Aldrich and used without further purification. The ligands 2-(1H-imidazol-2-yl)pyridine (pyim),^{62,63} 1-methyl-2-pyridin-2-yl-1H-benzimidazole (pyMebim),⁶⁴ and 1-methyl-2-(4'-thiazolyl)benzimidazole (Metbz) were prepared according to literature procedures or via adapted literature protocols. Deuterated solvents were obtained from SDS (CD₃SOC(D₃)) and Carlo Erba or Euriso-top (CDCl₃). Acetonitrile (HPLC, Prolabo) and tetrabutylammonium hexafluorophosphate ([ⁿBu₄N][PF₆], Acros) were used to prepare solutions for electrochemical measurements.

Synthesis method and complexes characterization. All synthetic manipulations were carried out under an atmosphere of dry, oxygen-free nitrogen using standard Schlenk techniques. The solvents were dried and distilled under nitrogen atmosphere before use. Elemental analyses were performed with a LECO CHNS-932 Elemental Microanalyzer (from Universidad Autónoma de Madrid). The analytical data for the new compounds were obtained from crystalline samples when possible. In some cases the data were reasonably accurate, but in others the agreement of calculated and found values for carbon was >0.4%, so that solvent molecules were introduced in the molecular formulae to improve agreement. IR spectra were recorded on a Jasco FT/IR-4200 spectrophotometer (4000–400 cm⁻¹ range) with Single Reflection ATR Measuring Attachment. FAB Mass spectra (position of the peaks in DA) were recorded with an Autospec spectrometer. The isotopic distribution of the heaviest set of peaks matched very closely that calculated for the formulation of the complex cation in every case. NMR samples were prepared under a nitrogen atmosphere by dissolving the suitable amount of compound in 0.5 mL of the respective oxygen-free deuterated solvent and the spectra were recorded at 298 K (unless otherwise stated) on a Varian Unity Inova-400 (399.94 MHz for ¹H; 161.9 MHz for ³¹P; 376.28 MHz for ¹⁹F; 100.6 MHz for ¹³C).

Typically, 1D ^1H NMR spectra were acquired with 32 scans into 32 k data points over a spectral width of 16 ppm. ^1H and $^{13}\text{C}\{^1\text{H}\}$ chemical shifts were internally referenced to TMS via the residual ^1H and ^{13}C signals of the corresponding solvents, CDCl_3 ($\delta = 7.26$ ppm and $\delta = 77.16$ ppm) and $(\text{CD}_3)_2\text{SO}$ ($\delta = 2.50$ ppm and $\delta = 39.52$ ppm), according to the values reported by Fulmer et al.⁶⁵ Chemical shift values are reported in ppm and coupling constants (J) in Hertz. The splitting of proton resonances in the reported ^1H NMR data is defined as s = singlet, d = doublet, t = triplet, st = pseudotriplet, q = quartet, sept = septet, m = multiplet, bs = broad singlet. 2D NMR spectra such as ^1H - ^1H gCOSY, ^1H - ^1H NOESY, ^1H - ^{13}C gHSQC and ^1H - ^{13}C gHMBC were recorded using standard pulse sequences. The probe temperature (± 1 K) was controlled by a standard unit calibrated with methanol as a reference. All NMR data processing was carried out using MestReNova version 10.0.2. Characterization data have been moved to the Supporting Information. TGA measurements were performed on a TGA Q50 V20.10 instrument from 25 to 800 °C in a flowing mixture of nitrogen/air (60:40 mL/min) at a heating rate of 10 °C·min⁻¹.

X-ray Crystallography. A summary of crystal data collection and refinement parameters for compounds [4][PF₆] and [5][PF₆] are given in Table S1. Single crystals of [4][PF₆] and [5][PF₆] were mounted on a glass fiber and transferred to a Bruker X8 APEX II CCD-based diffractometer equipped with a graphite monochromated MoK α radiation source ($\lambda = 0.71073$ Å). The highly redundant datasets were integrated using SAINT⁶⁶ and corrected for Lorentz and polarization effects. The absorption correction was based on the function fitting to the empirical transmission surface as sampled by multiple equivalent measurements with the program SADABS.⁶⁷

The software package SHELXTL (version 6.10)⁶⁸ was used for space group determination, structure solution, and refinement by full-matrix least-squares methods based on F^2 . A successful solution by direct methods provided most non-hydrogen atoms from the E-map. The remaining non-hydrogen atoms were located in an alternating series of least-squares cycles and difference Fourier maps. All non-hydrogen atoms were refined with anisotropic displacement coefficients. Hydrogen atoms were placed using a “riding model” and included in the refinement at calculated positions. CCDC reference numbers for [4][PF₆] and [5][PF₆] are 1519073 and 1519074 respectively.

The thermal ellipsoid molecular structures depicted in Figure 2 were plotted using the ORTEP-III program for crystal structure illustrations.⁶⁹

Electrochemical measurements. Electrochemical measurements were performed using a customized SPELEC (DropSens) equipment, a commercial fully integrated synchronized spectroelectrochemical device, that includes a bipotentiostat/galvanostat controlled by DropView (DropSens). All experiments were carried out using a three-electrode cell using a glassy carbon-disc with a diameter of 3 mm as working electrode, a platinum-wire as auxiliary electrode, and a silver-wire pseudo-reference electrode. Oxygen was removed from the solution by bubbling argon during 10 minutes and keeping the current of argon along the whole experiment. The formal potentials were determined by cyclic voltammetry (CV), scanning the potential at a scan rate of 0.15 V s⁻¹. All potentials were referred to the redox pair ferrocene/ferrocenium (Fc/Fc⁺)

took as internal standard. All voltammetric experiments were started and finished at a potential of -0.46 V and performed in a clockwise direction. Acetonitrile solutions of the complexes (10⁻³ M) were used in the presence of 0.1 M [ⁿBu₄N][PF₆] as supporting electrolyte.

Photophysical characterization. The UV-Vis absorbance spectra were measured with a Secomam XT5 UVi Light spectrometer. Quantum yields and PL emission spectra were recorded using a Hamamatsu absolute quantum yield C9920, whereas excited-state lifetimes were determined using a Hamamatsu Compact fluorescence lifetime spectrometer C11367 (Quantaurs-Tau).

Computational details. Density functional calculations (DFT) were carried out with the D.01 revision of the Gaussian 09 program package.⁷⁰ The Becke's three-parameter B3LYP exchange-correlation functional^{71,72} was used, together with the 6-31G** basis set for C, H, N, and S,⁷³ and the “double- ζ ” quality LANL2DZ basis set for the Ir atom.⁷⁴ Relativistic effects were accounted for by means of an effective core potential (ECP), which was used to replace the inner core electrons of Ir. The geometries of both the singlet ground electronic state (S_0) and the lowest-energy triplet state (T_1) were fully optimized for cations [1]⁺ to [5]⁺. In addition, the radical neutral species [1][•]-[5][•] resulting from the reduction of [1]⁺-[5]⁺ were also optimized in a doublet electronic state configuration (D_0). No symmetry restrictions were imposed. The geometries of [1][•]-[5][•] in the T_1 state and of [1][•]-[5][•] in the D_0 state were calculated at the spin-unrestricted UB3LYP level using spin multiplicity of three and two, respectively. All the calculations were performed in the presence of the solvent (acetonitrile). Solvent effects were considered within the self-consistent reaction field (SCRF) theory using the polarized continuum model (PCM) approach.^{75,76,77} Time-dependent DFT (TD-DFT)^{78,79,80} calculations of the lowest-lying 5 triplet states of the complexes were performed in the presence of the solvent at the minimum-energy geometry optimized for the ground state. The geometries of the first and second triplet excited states of [1][•] were first optimized at the TD-DFT level, and then reoptimized at the UB3LYP level to compare with the results obtained for T_1 from DFT calculations for all the other complexes.

Device preparation. LECs were prepared on top of a patterned indium tin oxide (ITO, 15 Ω square⁻¹) coated glass substrate (www.naranjosubstrates.com) previously cleaned as follows: a) sonication with soap, b) deionized water, c) isopropanol, and d) UV-O₃ lamp for 20 min. The thickness of the films was determined with an Ambios XP-1 profilometer. Prior to the deposition of the emitting layer, 80 nm of poly(3,4-ethylenedioxythiophene):poly-(styrenesulfonate) (PEDOT:PSS) (CLEVIOS™ P VP AI 4083, aqueous dispersion, 1.3–1.7% solid content, Heraeus) was coated in order to avoid the formation of pinholes and to increase the reproducibility of the cells. The emitting layer (100 nm) was prepared by spin-coating of a MeCN solution consisting of the iTMC with the addition of the ionic liquid (IL) 1-butyl-3-methylimidazolium hexafluorophosphate [Bmim][PF₆] (> 98.5%, Sigma-Aldrich) in a 4:1 molar ratio (iTMC:IL). The devices were then transferred to an inert atmosphere glovebox (< 0.1 ppm O₂ and H₂O, MBraun) and annealed during 1 hour at 100 °C. Finally, a layer (70 nm) of aluminium (the top electrode) was thermally evaporated onto the devices using an Edwards Auto500 evaporator integrated in the inert atmos-

phere glovebox. The area of the device was 6.5 mm². The devices were not encapsulated and were characterized inside the glovebox at room temperature.

Device characterization. The device lifetime was measured by applying a pulsed current and monitoring the voltage and the luminance versus time by a True Colour Sensor MAZeT (MTCSiCT Sensor) with a Botest OLT OLED Lifetime-Test System. The average current density is determined by multiplying the peak current density by the time-on time and dividing by the total cycle time. The average luminance is directly obtained by taking the average of the obtained photodiode

results and correlating it to the value of a luminance meter. The current efficiency is obtained by dividing the average luminance by the average current density. The electroluminescent (EL) spectra were measured using an Avantes AvaSpec-2048 Fiber Optic Spectrometer during device lifetime measurement.

Supporting Information. Synthesis of cationic iridium complexes, 2D 1H-1H NOESY spectra, TGA measurements, X-ray crystal structures, excited-state lifetimes, and theoretical calculations.

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