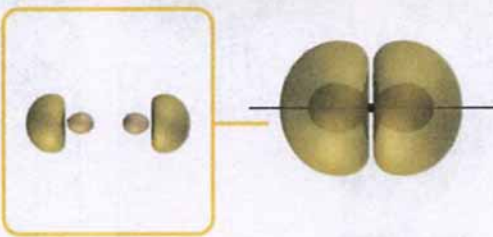


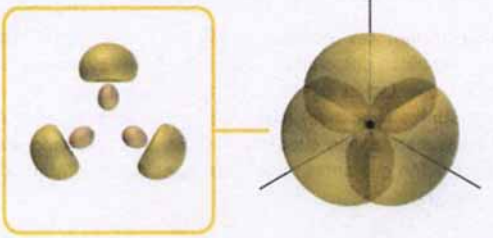

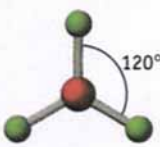
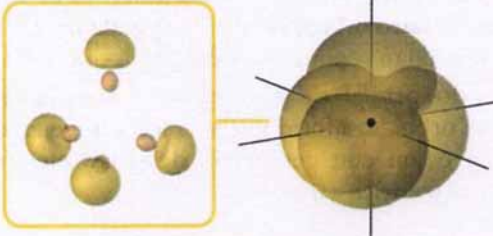

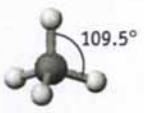
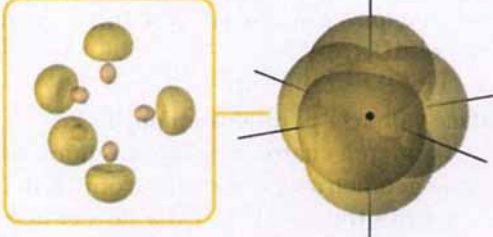
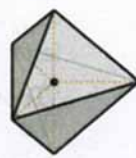
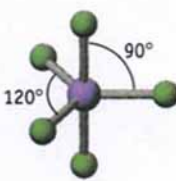
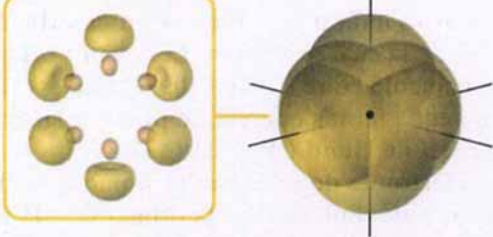

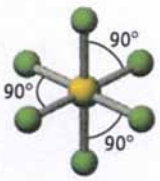


Distintos tipos de hibridaciones

Kotz, J. C.; Treichel, P. M., "Química y Reactividad Química", 5ª Ed., Thomson Paraninfo, 2003, pp 385.

Ordenamiento de orbitales híbridos	Figura geométrica	Ejemplo
<p>Dos pares de electrones sp</p> 	 <p>Lineal</p>	 <p>180° BeCl₂</p>
<p>Tres pares de electrones sp^2</p> 	 <p>Plana trigonal</p>	 <p>120° BF₃</p>
<p>Cuatro pares de electrones sp^3</p> 	 <p>Tetraédrica</p>	 <p>109.5° CH₄</p>
<p>Cinco pares de electrones sp^3d</p> 	 <p>Bipirámide trigonal</p>	 <p>90° 120° PF₅</p>
<p>Seis pares de electrones sp^3d^2</p> 	 <p>Octaédrica</p>	 <p>90° 90° 90° SF₆</p>



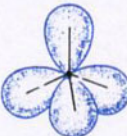

Douglas, B.; McDaniel, D.; Alexander, J., "Concepts and Models of Inorganic Chemistry", 3ª Ed., John Wiley & Sons, 1994, pp 80.

Douglas, B.; McDaniel, D.; Alexander, J., "Conceptos y Modelos en Química Inorgánica", 1ª Ed., Reverté, 1977, pp 83.

<i>Número de pares σ y pares solitarios</i>	<i>Configuración (considerando los pares solitarios dirigidos como otro grupo cualquiera)</i>	<i>Hibridación</i>
2	Lineal	sp
3	Trigonal plana	sp^2
4	Tetraédrica	sp^3
5	Bipirámide trigonal	$d_{z^2}sp^3$
5	Piramidal cuadrada	$d_{x^2-y^2}sp^3$
6	Octaédrica	d^2sp^3
7	Bipirámide pentagonal	d^3sp^3

Gillespie, R. J.; Humphreys, D. A.; Baird, N. C.; Robinson, E. A., "Química", Vol I, Reverté, 1990, pp 351.

Tabla 8.4 Orbitales híbridos o de enlace.

TIPO DE MOLÉCULA	ORBITAL HÍBRIDO O DE ENLACE	DISPOSICIÓN DE LOS ORBITALES	
AX_2	sp	Lineal	
AX_3	sp^2	Triangular	
AX_4	sp^3	Tetraédrica	
AX_5	sp^3d	Bipiramidal trigonal	
AX_6	sp^3d^2	Octaédrica	